**Which Nitrate will produce the most Precipitate? (Lab)**

**Problems:**

1. **What are the solubility rules?**
2. **How can you make 50 ml of a .1 M solution of Cu(NO3)2, Mg(NO3)2 and Al(NO3)3?**
3. **How can you make 50 ml of a .1 M Potassium Chromate (polyatomic: look up)?**
4. **If you mixed the 10 ml of the 3 nitrates with 10 ml of the .1 M Potassium Chlorate, which will create the most precipitate?**

**Material: Solid form of the chemicals above, distilled water, funnel and filter paper.**

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| --- | --- | --- | --- | --- |
|  | **Molar** **Mass** | **Mass needed for 50 ml of a .1 M sol.** | **Will it make a precipitate?** | **Precip. Collected.** |
| **Potassium****Chromate**  |  |  |  |  |
| **Mg(NO3)2** |  |  |  |  |
| **Al(NO3)3** |  |  |  |  |
| **Cu(NO3)2** |  |  |  |  |

**Conclusion: 50 words.**