**Unit 6 prob VI**

1. **If it took1 5.98 ml of 12.5 M NaOH to neutralize 279.45 ml of an unknown Hydroflouric acid (HF), what is the molarity of the acid?**
2. **If it took 5.5 ml of an unknown sulfuric acid (H2 SO4) to neutralize 362.5 ml of a 2.5 M calcium hydroxide, what is the molarity of the acid?**
3. **If it took 35.98 ml of .5 M Mg(OH)2 to neutralize 49.45 ml of an unknown Hydroflouric acid, what is the molarity of the acid?**
4. **If it took 55.5 ml of an unknown Nitric acid (HNO3) to neutralize 524.5 ml of a 2.59 M Lead (III) hydroxide, what is the molarity of the acid?**
5. **How many grams of NaCl would you need to dissolve in one liter of water to make a 1.13 molar solution (Hint: Molar mass x .13)?**
6. **If 20 ml of a 2.5 M KOH solution was used to neutralize 32 ml of HNO3, what is the molarity of the acid (HNO3 )?**
7. **If 45 ml of HCl was used to neutralize 20.1234 ml of a .5 M Ca(OH)2 solution, what is the molarity of the acid?**
8. **If you wanted to make a .1 M solution of cesium fluoride, how many grams would you dissolve in 1 liter of water? (same as #5)**