**Titration Lab III**

**Problems:**

1. **How can you neutralize 10 ml of 1.5 M HCl (or HNO3) and 10 ml of a 1.5 M H2SO4?**
2. **How do you think the volumes of NaOH required will compare?**
3. **Which acid will produce the most H3O ions?**

**Material: HCl, H2SO4, graduated buret, PT indicator, E. Flask.**

**Data:**

|  |  |
| --- | --- |
|  | **NaOH (ml)** |
| **HNO3** |  |
| **H2SO4** |  |

**Conclusion: 50 words. Answer the following:**

1. **Which acid used the most base?**
2. **What are hydronium ions?**
3. **How do acids create hydronium ions?**
4. **What determines the strength of an acid?**
5. **If the pH’s of the acids were the same how could one acid create more H3O ions?**
6. **Which acid had twice as many hydrogens in it? How could this affect its strength?**