**The Mole and conversions.**

**When ever you are converting 1st start by putting the given measurement over 1 to begin dimensional analysis.**

**Example: convert 10 g of Mg to moles.**

**10 g of Mg**

**1**

**When converting from grams to Moles or vice versa, you can use the molar mass (P. Table) to make your conversion factor.**

**10 g of Mg x 1 Mole Mg = .411 Moles of Mg**

1. **24.305 g of Mg**

**When you are converting to grams to pieces (or moles) you need to use the number that represents the Mole.**

**Example: How many of the smallest possible pieces (atoms, molecules or formula units) are in 25.2 g of water?**

**25.2 g H2O x 1 Mole H2O x 6.02 x 1023 molecules=**

**1 18 g H2O 1 Mole H2O**

**8.43 x 1023 molecules of H2O**

**Turn in the following:**

1. **Convert 3.4 g of N2O3 to moles.**
2. **Convert 12.9 g of AlCl3 to formula units.**
3. **Convert 5.4 moles of Ammonium Acetate to grams.**