**Quantum Numbers**

**Quantum numbers: Are in all electron configurations and are used to locate the specific location of electrons in an atom.**

**There are four different kinds of quantum numbers.**

1.)Principle number ( *n* ): indicates the energy level that the electron is in.

**example: 1s2**

**2.) Angular momentum number (l ): indicates the types the type of orbital that the electron is in (s, p, d or f).**

**3.) Magnetic quantum number (Ml ): is the number of possible paths for the electrons to be traveling on.**

**4.) Electron spin number (Ms): is the specific direction of the orbiting electrons. Electron spin is always +1/2 or –1/2.**