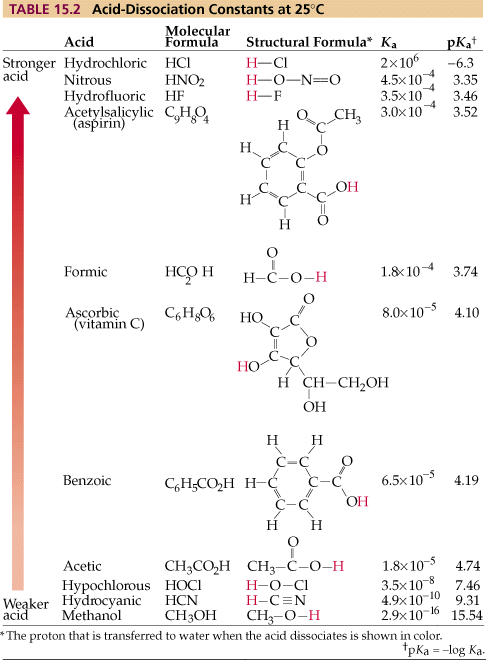
**Ka and pKa**

**Values fro Ka and pKa (-log Ka) for some typical weak acids and 1 strong are:**



**Note that pKa decreases as Ka increases. As indicated in the equilibrium equation (previous notes), the larger the value of Ka, the stronger the acid.**

**Thus, methanol (Ka = 2.9 x 10-16, pka = 15.54)is the weakest acid listed on the given table.**

**Nitrous acid (Ka = 4.5 x 10-4, pKa = 3.35) is the strongest of the weak acids.**

**Strong acids, such as HCl, have Ka values that are much greater than 1 and pKa values that are negative.**