**Law of Definite Proportions.**

**Learning Target:**

**The students will learn what the Law of Definite Proportions is and how it is related to chemical reactions.**

**Law of Definite proportions**

**Law of Definite Proportions: the subscripts of a compound can only be changed if it goes through a chemical change.**

**\*\*If the subscripts are changed that means that the compounds is completely different.**

**Example:**

**CO2 (Carbon Dioxide): Required for life.**

**CO (Carbon Monoxide): kills plants and animals.**

**\*\*This is why you can’t change the subscripts when you are balancing chemical equations.**

**Example:**

**H2 + O2 🡪 H2O**

**\*\*Accorinding to the Conservation of mass all chemical equations must balance. According to the Law of Definite Proportions, when balancing a chemical equation you need to change the Coefficients not the Subscripts.**