**CuSO4 + H2O Lab**

**Problems:**

1. **How can you boil off the water that is physically bonded to a Copper Sulfate (Blue powder)?**
2. **How can you calculate the water that is lost from the compound?**
3. **How can you calculate the % of water lost from the compound?**

**Material: Copper Sulfate, tin foil, hot plate, balance.**

**Data:**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Mass of foil** | **Mass before heat – foil** | **Mass after heat- foil** | **Difference** | **% water lost** |
|  |  |  |  |  |

**Heat the powder until it turns very light blue to white in color. Try not to let it turn black.**

**% water lost = difference / mass before the heat x 100**

**Calculate % error using the actual percentage which is: 36.6 %**

**% error = actual – experimental/actual x 100**

**Conclusion: Answer ?’s From Notes.**