Conductivity Lab.

**Problem (s):**

1. **Which type of reaction will change the conductivity of an acid the most, neutralization or an exothermic reaction?**
2. **Will the conductivity of the acid go up, down or stay the same as you add the base to it?**

**Material: Power source, electrolysis leads, acid, base, beaker (s), Metal and distilled water.**

**Data:**

|  |  |  |  |
| --- | --- | --- | --- |
|  | **Obser. w/ electricity** |  | **Observation w/ electricity** |
| **Acid** |  | **H2O** |  |
| **Acid + Mg** |  | **X** | **X**  |
| **Acid + base** |  | **X**  | **X** |

**Conclusion.**

1. **Use complete sentences**
2. **Explain how you solved the problems using the following terms:**

**Conductivity, Reaction rate, Exothermic, Neutralization, Control and Variables**