**Chemical Systems / Enthalpy Lab**

**Problems:**

1. **How will the change in temperature of 50 ml of water compare when you:**
2. **Put heated Penny in it?**
3. **Put heated Nickel in it?**
4. **How can you calculate the heat released from the metals to the water?**

**Material: B. burner, tongs, Penny, Nickel, beaker.**

**Data:**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
|  | **Mass of 20 ml H2O** | **S.H. of H2O** | **IT H2O** | **FT**  **H2O** | **CT**  **H2O** | **C. in heat of H2O** |
| **Penny** | **20 g** | **4.184** |  |  |  |  |
| **Nickel** | **20g** | **4.184** |  |  |  |  |

**Calculations: Heat = mass x SHx CT**

**Conclusion: From Notes.**