Catalysts, Reaction Rates

**& Salts.**

Reaction rate: is the speed at which a reaction takes place.

**Reaction rate can be determined by timing how fast the reactants of a reaction are used.**

**Catalyst: is a substance that speeds up the reaction rate without chemically reacting with the things involved.**

**Catalyst work by lowering the amount of energy required (called: activation energy) for the reaction to occur.**

**Example: Mg + acid**

**no cata. Catalyst present**

**Slower reaction Faster reaction**

**Mg**

**Salts: are ionic compounds that are direct result of a neutralization reaction.**

**Neutralization: is a reaction between an acid and a base that is the result in a pH change closer to 7.0.**

**Example:**

**2KOH + H2SO4🡪 2H2O + K2SO4**

**Base Acid Water Salt**

**\*\*Water and some kind of salt are always the products of a “real” neutralization.**