**Yellow and Blue Precipitate**

**Lab.**

**Problems:**

1. **How can you perform 4 flames tests on 4 different compounds?**
2. **How can you observe what happens when you add acid to the same 4 compounds?**
3. **What do you think will happen when you mix 5ml of potassium carbonate with 5 ml of copper (II) sulfate?**
4. **What do you think will happen when you mix 5 ml of sodium iodide with 5 ml of Lead (II) Nitrate?**
5. **How can you run the same tests the next day to determine what the Blue and yellow precipitates are?**

**Materials: see data table.**

**Data:**

|  |  |  |
| --- | --- | --- |
|  | **Flame test**  **Before Mixing** | **Acid Test**  **Before**  **Mixing** |
| **K2CO3** |  |  |
| **CuSO4** |  |  |
| **NaI** |  |  |
| **Pb(NO3)2** |  |  |

|  |  |  |  |
| --- | --- | --- | --- |
|  | Flame test | Acid test  (Bubbles mean CO3 is present) | Formula |
| Yellow Precipitate |  |  |  |
| Blue Precipitate |  |  |  |

Yellow Ion exchange:

Blue Ion exchange:

**Conclusion. Use complete sentences to answer the following:**

1. **What are flame tests for? How did you use them in this lab.?**
2. **What was the formula of the Yellow Precipitate? Blue Precipitate?**
3. **What did the acid test indicate?**
4. **How did the acid test help you identify one of the products?**
5. **What reactants did you use to create the blue precipitate?**
6. **What reactants did you use to create the yellow precipitate?**