**Assigning Oxidations and summing charges**

**Within Polyatomic Ions.**

**If you are remember the formula and the oxidation of a Polyatomic Ion you can get the oxidation charges of the elements within it.**

**Example: Chlorate. ClO4-1.**

 **? -2**

**O is -2 because it isn’t with a peroxide.**

**ClO4’s over all charge is -1.**

**So:**

**1(?) + 4(-2) = -1**

**? + (-8) = - 1**

**Cl O’s ClO4-1**

**The sum of Cl’s oxidation is +7.**