Activity Series Lab III

Problems:

1. How can you convert a nugget of zinc into moles?
2. How can you convert 1 cm of Mg into moles?
3. How can you use an activity series (on the back of your periodic table) to predict if these two metals will react with silver nitrate and copper (II) sulfate?

Material: scale, Zn, Mg, Silver Nitrate (5 ml), Copper (II) Sulfate, test tubes.

Data:

Mass of Zn+2:\_\_\_\_\_\_\_\_\_\_\_\_

Calculation of moles of Zinc:

Mass of Mg:\_\_\_\_\_\_\_\_\_\_\_\_\_

Calculation of moles of Mg:

|  |  |  |  |
| --- | --- | --- | --- |
| Reactants | Prediction(Yes or No) | Obser. | Products |
| Mg + AgNO3 | Yes |  |  |
| Mg+CuSO4 |  |  |  |
| +2Zn + AgNO3 |  |  |  |
| +2Zn + CuSO4 |  |  |  |

Ion exchanges (4):

Mg +\_\_ AgNO3🡪 \_\_Mg (NO3)2 +\_\_ Ag

Conclusion: Notes